

Crystal structure of a new compound $\text{Er}_2\text{MnW}_{0.8}\text{O}_{6.4}$ phase

G. Chen^{a,b,*}, Y. Imagawa^a, H. Satoh^a, H. Nakano^c, N. Kamegashira^a

^a Department of Materials Science, Toyohashi University of Technology, Tempaku-cho, Toyohashi 441-8580, Japan

^b Department of Applied Chemistry, Harbin Institute of Technology, 92 West Dazhi Street, Harbin 150001, PR China

^c Department of Electrical and Electronics Engineering, Kokushikan University, Setagaya, Setagaya-ku 154-8515, Japan

Received 30 July 2004; received in revised form 7 November 2004; accepted 15 December 2004

Available online 12 July 2005

Abstract

A new compound of $\text{Er}_2\text{MnW}_{0.8}\text{O}_{6.4}$ system was synthesized by solid-state reaction method under low oxygen partial pressure. A single phase with a composition $\text{Er}_2\text{MnW}_{0.8}\text{O}_{6.4}$ was obtained after trying various composition ratios and its crystal structure was refined by powder X-ray diffraction data using the Rietveld analysis. The space group was $P3_121$ belonging to trigonal system with the lattice parameters of $a = 0.725$ and $c = 0.936$ nm. The electron diffraction data also showed this size of the unit cell. The derived structure was not a pyrochlore-related structure because no systematic stacking layers by hexagonal tungsten bronze type appeared.

© 2005 Elsevier B.V. All rights reserved.

Keywords: Structural analysis; Compounds; X-ray diffraction; Pyrochlore-related structure; Rare earth

1. Introduction

Complex oxides $\text{Ln}_2\text{B}_{2-x}\text{B}'_x\text{O}_7$ phases with a pyrochlore and its related structure have been studied from interesting physical and chemical properties by various combinations of rare earth (Ln) and transition metal (B, B') ions. Among them Mn-containing compounds $\text{Ln}_2\text{Mn}_2\text{O}_7$ can be prepared only under high pressure [1–3], but some compounds such as $\text{Ln}_2\text{Mn}_{1-x}\text{M}_x\text{O}_7$ (M = W, Mo, etc.) were synthesized under normal conditions [4–11]. In these Mn-containing oxides, the oxidation states of Mn ions are known to change under their varying synthesizing atmospheres. Bazuev et al. [4–6] and Chen et al. [7–9] have shown the possibility of synthesizing mixed-pyrochlores, $\text{Ln}_2\text{Mn}_x\text{M}_{2-x}\text{O}_7$ (Ln = Er or Ho, M = Ta, Nb or Mo), under an ordinary pressure and in various atmospheres. X-ray diffraction (XRD) data have shown that the compounds possess rhombohedral or hexagonal (trigonal) structures.

In recent years, some of these oxides have been investigated with XRD-Rietveld analysis and high-resolution

transmission electron microscopy (HRTEM) [12–14]. From the results of structural analysis $\text{Ho}_2\text{Mn}_{2-x}\text{Ta}_x\text{O}_7$ [13] and $\text{Er}_2\text{Mn}_{2/3}\text{Mo}_{4/3}\text{O}_7$ [14] have a trigonal symmetry (zirkelite-type) and a monoclinic symmetry (zirconolite-type), respectively. This difference in the structures was caused by the stacking sequence along the c -direction of a pair of A cation and B cation layers; the B cations formed a hexagonal tungsten bronze (HTB)-type layer that is a characteristic frame-layer in the pyrochlore-related structure [15,16].

In the present study, a new oxide $\text{Er}_2\text{MnW}_{0.8}\text{O}_{6.4}$ was synthesized in a reducing atmosphere at 1623 K. Its crystal structure was characterized by X-ray diffraction data using the Rietveld analysis [17,18] and HRTEM.

2. Experimental

The starting materials were Er_2O_3 , Mn_2O_3 and WO_3 with a purity of 99.99% (Rare Metallic Co. Ltd.). In order to adjust the stoichiometry of the starting materials Er_2O_3 was prepared at 1273 K for 24 h in an Ar atmosphere, WO_3 was heated at 673 K for 12 h in air and Mn_2O_3 at 1073 K for 72 h in air and then quenched to room temperature.

* Corresponding author. Tel.: +86 451 86413753; fax: +86 451 86413707.
E-mail address: gchen@hit.edu.cn (G. Chen).

These starting materials were thoroughly mixed in various molar ratios to obtain a single phase of Er–Mn–W–O system. The mixtures were pressed into pellets and heated at 1623 K for 48 h in Ar containing 1% H₂. The products were identified with a powder X-ray diffractometer using Cu K α radiation, and the experimental data were further analyzed by the Rietveld method. The specimen of TEM was crushed in a mortar and scooped with carbon-coated microgrids on a copper mesh. Then the crushed specimen was investigated with a JEM-2000EX TEM operating at 200 kV. High-resolution structure imaging was performed by JEM-3000F TEM operating at 300 kV with spherical aberration of $C_s = 1.0$ nm.

3. Results and discussion

When the syntheses were done with the molar ratio Mn:W = 1:1, Er₂MnWO₇ with a pyrochlore-related structure was not obtained as a single phase because ErMnO₃ was also found as the second phase. However, it is shown that Er₂MnW_{0.8}O_{6.4} synthesized by the molar ratio Mn:W = 1:0.8 was most similar to a single phase since only small quantity of MnO was found as the second phase. For Mn:W > 1:0.8, more much quantity of MnO was found, while Mn:W < 1:0.8 ErMnO₃ phase was found.

The X-ray diffraction patterns of Er₂MnW_{0.8}O_{6.4} are similar to those of Ho₂MnTaO₇ [13] having a pyrochlore-related structure as is shown in Fig. 1. The diffraction peaks are indexed based on a hexagonal cell, but the cell lattice parameters were $a = 0.7252$, $c = 0.9369$ nm, which was half the size of the c -dimension in Ho₂MnTaO₇.

In order to investigate the structure of Er₂MnW_{0.8}O_{6.4}, a Rietveld analysis was carried out by the XRD data based on two structure models, the type 1 and the type 2, with the space group $P3_121$. Tables 1 and 2 list the refined atomic coordinates. From these data it was shown that the type 2 (the value of the lattice parameter c was one half the type 1) gives better results than those of the type 1. The reliability factor R_{wp} of structural refinement had 14.33% for the type 1, while R_{wp} had the smaller 13.12% if the type 2 was used. The XRD patterns calculated with the structural

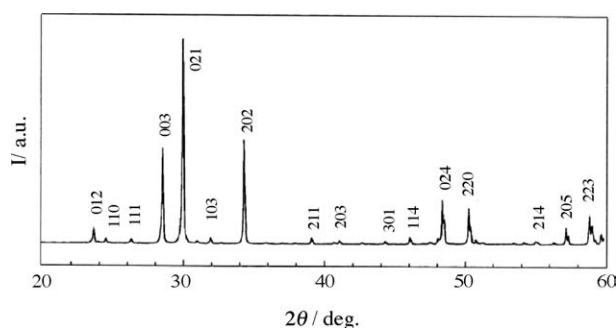


Fig. 1. Powder XRD pattern of Er₂MnW_{0.8}O_{6.4}. The peaks are indexed as a hexagonal cell with lattice parameters of $a = 0.7242$ and $c = 0.9365$ nm.

Table 1
Atomic coordinates of Er₂MnW_{0.8}O_{6.4} (type 1)

Atom	Site	Atomic coordinates		
		x^a	y^a	z^a
Er(1)	3(a)	0.8779(9)	0	1/3
Er(2)	3(a)	0.3341(4)	0	1/3
Er(3)	6(c)	0.6621(9)	0.1586(1)	−0.0102(5)
Mn, W(1)	3(b)	0.8685(3)	0	5/6
Mn, W(2)	3(b)	0.3254(2)	0	5/6
Mn, W(3)	6(c)	0.5166(3)	0.3424(2)	0.1577(1)
O(1)	6(c)	0.2035(2)	0.2283(1)	0.1577(1)
O(2)	6(c)	0.5633(1)	0.6084(1)	0.1955(2)
O(3)	6(c)	0.1955(1)	0.6400(3)	0.1464(1)
O(4)	6(c)	−0.0493(1)	0.3164(1)	0.0547(1)
O(5)	6(c)	0.0363(2)	0.8235(2)	0.0547(1)
O(6)	6(c)	0.560(3)	0.4000(1)	0.0565(5)
O(7)	6(c)	0.555(5)	0.8235(1)	0.0591(1)

Space group: $P3_121$; $a = 0.7252(2)$ nm, $c = 1.8739(9)$ nm. $R_{wp} = 14.33\%$, $R_p = 10.23\%$. $R_E = 5.17\%$. $R_1 = 6.62\%$. $R_F = 6.44\%$. Site occupancies of (Mn,W(1)), (Mn,W(2)) and (Mn,W(3)) are 0.9, 0.9 and 1.0, respectively.

^a The standard deviation of the last digit is given in parenthesis.

model of the type 2 were in agreement with the observed ones.

The crystal structure of Er₂MnW_{0.8}O_{6.4} based on the structural model of the type 1 and 2 is shown in Fig. 2. In the case of type 1, the crystal structure is similar to Ho₂MnTaO₇ [13] in which A and B cation layers alternately form a pair of layers along the c -direction. On the other hand, in the case of type 2, no pair was observed in the structure, and accordingly, the c -dimension was up to half the size of that in pyrochlore-related compounds. The results indicated that this structure in Er₂MnW_{0.8}O_{6.4} is not a pyrochlore-related structure. In addition, the B cations did not come to form a characteristic layer of an HTB-type layer [15,16].

The results of the structural refinement for the type 2 shown that there are two kinds of oxygen octahedra surrounding the central (Mn, W) ions where manganese and tungsten ions are randomly distributed over the B-site. Each of these B-site ions is coordinated to six oxygen ions, but the

Table 2
Atomic coordinates of Er₂MnW_{0.8}O_{6.4} (type 2)

Atom	Site	Atomic coordinates		
		x^a	y^a	z^a
Er(1)	6(c)	0.653(4)	0.147(9)	0.020(2)
Mn, W(1)	3(a)	0.848(4)	0	1/3
Mn, W(2)	3(a)	0.325(9)	0	1/3
O(1)	3(a)	0.583(1)	0	1/3
O(2)	3(b)	0.152(3)	0	1/3
O(3)	3(b)	0.538(2)	0	5/6
O(4)	6(c)	0.991(2)	0.349(5)	−0.067(5)
O(5)	6(c)	0.756(5)	0.969(5)	0.122(7)

Space group: $P3_121$; $a = 0.7252(1)$ nm, $c = 0.9369(8)$ nm. $R_{wp} = 13.12\%$, $R_p = 9.86\%$. $R_E = 5.18\%$. $R_1 = 6.93\%$. $R_F = 6.65\%$. Site occupancies of (Mn,W(1)) and (Mn,W(2)) are 1.0 and 0.8, respectively.

^a The standard deviation of the last digit is given in parenthesis.

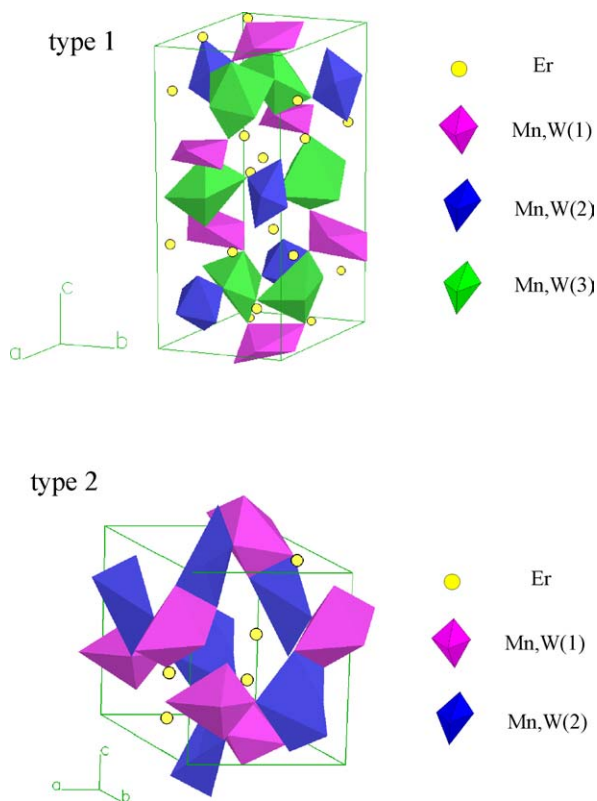


Fig. 2. The crystal structure of $\text{Er}_2\text{MnW}_{0.8}\text{O}_{6.4}$ based on the structural model of the type 1 and 2.

shape of these MO_6 field in both types are much distorted from the regular octahedron. Their stacking mode is apparently similar to those of the pyrochlore-related structures such as pyrochlore, zirkelite and zirconolite, but no regularity in the stacking of MO_6 characteristic of these structures was seen.

Fig. 3 shows the selected area diffraction (SAD) patterns obtained from the $[\bar{1}00]$ in (a), $[\bar{1}10]$ in (b) and $[001]$ zone axes in (c). All of the reflections could be indexed based on the hexagonal cell determined by XRD. On the $(001)^*$ plane, the pattern was same as that of $\text{Ho}_2\text{MnTaO}_7$ with space group of $P3_121$ [13]. However, along the c^* -

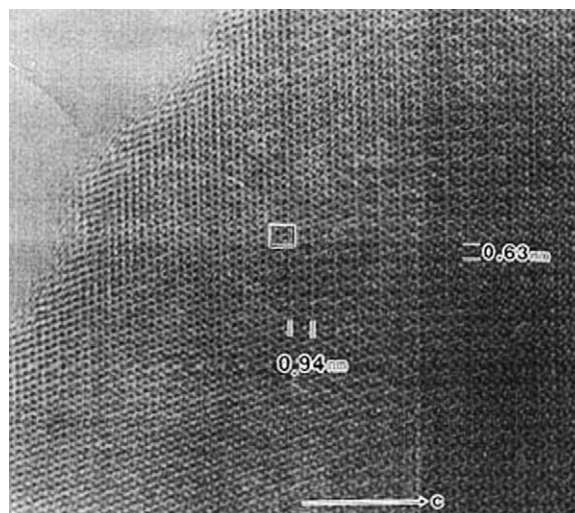


Fig. 4. HRTEM image of $\text{Er}_2\text{MnW}_{0.8}\text{O}_{6.4}$ taken from the $[100]$ zone axis.

axis, the cell size was calculated to be half the size of the c -dimension in $\text{Ho}_2\text{MnTaO}_7$. In Fig. 3(a), intensity of 021 reflection was different from that of $0\bar{2}1$ reflection, and the fact indicated that the cell was a trigonal not a hexagonal.

Fig. 4 shows a HRTEM image of $\text{Er}_2\text{MnW}_{0.8}\text{O}_{6.4}$ taken from the $[100]$ zone axis. The lattice fringes with repeat distance of 0.94 nm was observed, where each fringe consisted of three cation layers and the repeat distance corresponded to the c -dimension. The lattice image was obviously different from that of pyrochlore-related structures, $\text{Ho}_2\text{MnTaO}_7$ [13] and $\text{Er}_2\text{Mn}_{2/3}\text{Mo}_{4/3}\text{O}_7$ [14], in which A and B cation layers arranged alternately form a pair of layers along the c -direction. In this case, no pair was observed in the structure and the c -dimension was up to half the size of that in pyrochlore-related compounds. These results were in agreement with those of XRD-Rietveld analysis. In Fig. 4, the weak repeat distance of 0.63 nm corresponded to (100) spacing. Black spots were believed to reflect cations in the thin crystal, and the unit cell was enclosed by the rectangle.

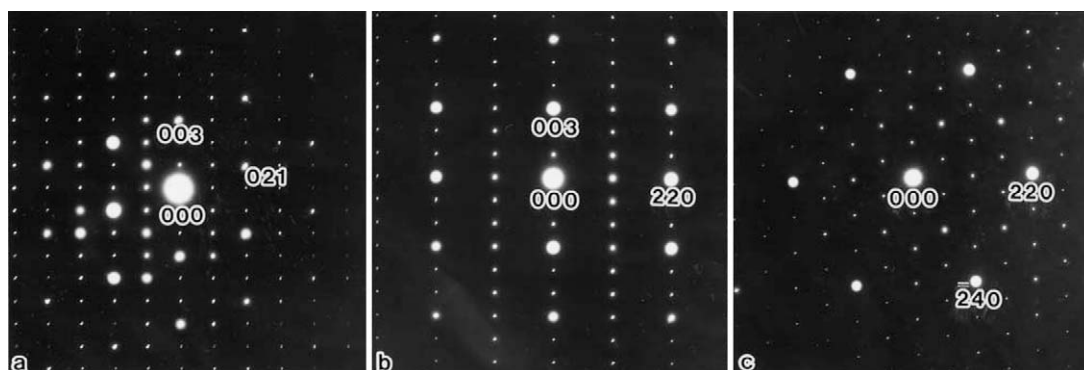


Fig. 3. Selected area diffraction patterns taken from the $[\bar{1}00]$, $[\bar{1}10]$ and $[001]$ zone axes.

4. Conclusions

The new compound, $\text{Er}_2\text{MnW}_{0.8}\text{O}_{6.4}$, synthesized from a mixture of Er_2O_3 , Mn_2O_3 and WO_3 was confirmed to have a trigonal cell. The cell parameters were determined to be $a=0.725$ and $c=0.936$ nm by Rietveld analysis and HRTEM image. The c -dimension was up to half the size of that in $\text{Ho}_2\text{MnTaO}_7$, which has a pyrochlore-related structure. TEM data revealed that the A-cations and B-cations did not form a pair of layers along the c -direction, and accordingly, it could be assumed that the B-cations did not form HTB-type layers as seen in pyrochlore-related structures. The same result was also proved from the Rietveld analysis of $\text{Er}_2\text{MnW}_{0.8}\text{O}_{6.4}$. Therefore, the structure in $\text{Er}_2\text{MnW}_{0.8}\text{O}_{6.4}$ is not a pyrochlore-related structure but a new type structure with $P3_121$ space group.

Acknowledgement

This work was supported by the Grant-in-Aid for Scientific Research (B) (No. 13450259) by the Japan Society for the Promotion of Science.

References

- [1] H. Fijinaka, N. Kinomura, M. Koizumi, Y. Miyamoto, S. Kume, Mater. Res. Bull. 14 (1979) 1113.
- [2] M.A. Subramanian, G. Aravamudan, G.V. Subba Rao, Mater. Res. Bull. 15 (1980) 1401.
- [3] M.A. Subramanian, C.C. Torardi, D.C. Johnson, J. Pannetier, A.W. Slight, J. Solid State Chem. 72 (1988) 24.
- [4] G.V. Bazuev, O.V. Makarova, G.P. Shveikin, Zh. Neog. Khim. 28 (1983) 1919.
- [5] G.V. Bazuev, O.V. Makarova, G.P. Shveikin, Zh. Neog. Khim. 30 (1985) 1253.
- [6] G.V. Bazuev, G.P. Shveikin, Zh. Neog. Khim. 32 (1987) 1764.
- [7] G. Chen, H. Takenoshita, N. Kamegashira, Mater. Chem. Phys. 46 (1996) 43.
- [8] G. Chen, H. Takenoshita, N. Kamegashira, J. Alloys Compd. 238 (1996) 28.
- [9] G. Chen, H. Takenoshita, J. Kondo, N. Kamegashira, J. Alloys Compd. 359 (2003) 209.
- [10] T. Shirakami, H. Nakano, K. Urabe, A. Matsuo, N. Kamegashira, Solid State Ionics 121 (1999) 85.
- [11] D. Ismunandar, B.J. Kennedy, B.A. Hunter, J. Alloys Compd. 302 (2000) 94.
- [12] H. Nakano, N. Kamegashira, Mater. Res. Bull. 35 (2000) 1631.
- [13] H. Nakano, N. Kamegashira, K. Urabe, J. Am. Ceram. Soc. 83 (2000) 2305.
- [14] H. Nakano, N. Kamegashira, J. Am. Ceram. Soc. 84 (2001) 1374.
- [15] F. Mazzi, R. Munno, Am. Mineral. 68 (1983) 262.
- [16] T.J. White, Am. Mineral. 69 (1984) 1156.
- [17] F. Izumi, J. Cryst. Soc. Jpn. 27 (1985) 23.
- [18] F. Izumi, J. Mineral. Soc. Jpn. 17 (1985) 37.